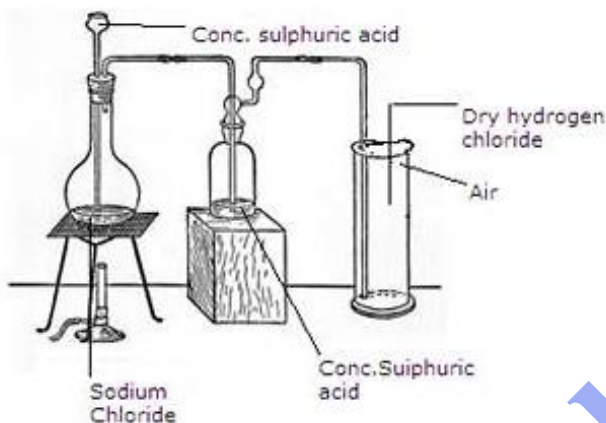


**EXERCISE****Question 1:**

Draw a labelled diagram for the laboratory preparation of hydrogen chloride gas. Give the balanced equation for the reaction.

**Solution 1:**

Labelled Diagram for laboratory preparation of Hydrogen chloride is:

**Question 2:**

Name the drying agents:

- used in drying hydrogen chloride gas.
- phosphorus pentoxide and calcium oxide are good drying agent but they cannot be used to dry hydrogen chloride gas. Why?

**Solution 2:**

- Hydrogen chloride is dried by passing through conc. Sulphuric acid.
- Phosphorous pentoxide and CaO cannot be used to dry HCl because they react with HCl.  
 $2\text{P}_2\text{O}_5 + 3\text{HCl} \rightarrow \text{POCl}_3 + 3\text{HPO}_3$   
 $\text{CaO} + 2\text{HCl} \rightarrow \text{CaCl}_2 + \text{H}_2\text{O}$

**Question 3:**

Explain why:

- (a) Anhydrous HCl is poor conductor while aqueous HCl is excellent conductor.
- (b) When the stopper of a bottle full of hydrogen chloride gas is opened there are fumes in the air.
- (c) A solution of hydrogen chloride in water turns blue litmus red, and conducts electricity, while a solution of the same gas in toluene:
  - (i) has no effect on litmus, and
  - (ii) does not conduct electricity.
- (d) thick white fumes are formed when a glass rod dipped in  $\text{NH}_2\text{OH}$  is brought near the mouth of a bottle full of HCl gas.
- (e) dry hydrogen chloride gas does not affect a dry strip of blue litmus paper but it turns red in the presence of a drop of water.
- (f) hydrogen chloride gas is not collected over water.

**Solution 3:**

- (a) Anhydrous HCl is poor conductor due to the absence of ions in it whereas aqueous HCl is excellent conductor since it contains ions.
- (b) When the stopper is opened HCl gas comes in contact with water vapors of air and gives white fumes due to the formation of hydrochloric acid.
- (c) A solution of HCl in water gives hydronium ions and conducts electricity, but HCl is also soluble in dry toluene, but in that case it neither (i) turns blue litmus red (ii) nor does conducts electricity. This indicates the absence of  $\text{H}^+$  ions in toluene showing thereby that hydrogen chloride is a covalent compound.
- (d) When ammonium hydroxide is brought near the mouth of HCl, dense white fumes are formed due to the formation of ammonium chloride.  
$$\text{HCl} + \text{NH}_4\text{OH} \rightarrow \text{NH}_4\text{Cl} + \text{H}_2\text{O}$$
- (e) Dry hydrogen chloride is not acidic whereas moist Hydrogen chloride is acidic. In presence of a drop of water HCl gas dissolves in water and forms hydrochloric acid which turns blue litmus paper red.
- (f) Hydrogen chloride is not collected over water as it is highly soluble in water.

**Question 4:**

Write the main difference in hydrogen chloride gas and hydrochloric acid.

**Solution 4:**

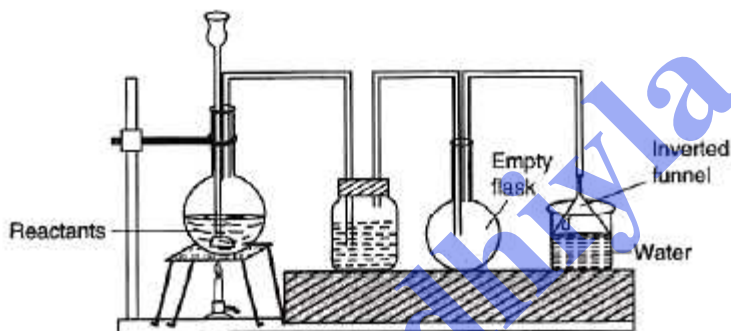
Difference between Hydrogen chloride gas and Hydrochloric acid is:

Hydrogen chloride gas	Hydrochloric acid
1. Dry hydrogen chloride gas does not turn blue litmus red due to non-acidic character. 2. Hydrogen chloride gas does not conduct electricity.	1. Being acidic it turns blue litmus red. 2. Hydrochloric acid is a good conductor of electricity.

**Question 5:**

The given set up in the figure is for the preparation of an acid.

(a) Name the acid prepared by this method.



(b) name the reactants used.

(c) why empty flask is used

(d) what is the drying agent used? Why is this drying agent chosen?

(e) what is the role of inverted funnel in the arrangement

**Solution 5:**

Hydrochloric acid is prepared by this method.

(a) The reactants are sodium chloride and Sulphuric acid.

(b) The empty flask acts as Anti-Suction device. In case the back suction occurs the water will collect in it and will not reach the generating flask.

(c) The drying agent is Conc. Sulphuric acid. Sulphuric acid is chosen as drying agent because it does not react with HCl.

(d) The Inverted funnel :

Prevents or minimizes back suction of water.

Provides a large surface area for absorption of HCl gas.

**Question 6:**

Write an equation for the reactions of aqueous hydrochloric acid on:

- (a) silver nitrate solution                      (b) magnesium foil  
(c) caustic soda solution                      (d) zinc carbonate  
(e) lead nitrate solution                      (f) copper oxide

**Solution 6:**

Equations are:

- (a)  $\text{AgNO}_3 + \text{HCl} \rightarrow \text{AgCl} + \text{HNO}_3$   
(b)  $\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$   
(c)  $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$   
(d)  $\text{ZnCO}_3 + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2\text{O} + \text{CO}_2$   
(e)  $\text{Pb}(\text{NO}_3)_2 + 2\text{HCl} \rightarrow \text{PbCl}_2 + 2\text{HNO}_3$   
(f)  $\text{CuO} + 2\text{HCl} \rightarrow \text{CuCl}_2 + \text{H}_2\text{O}$

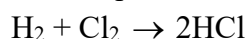
**Question 7:**

- (a) Name an element which reacts with hydrogen to form a compound which is strongly acidic in water.  
(b) Explain why dilute hydrochloric acid cannot be concentrated by boiling beyond 22.2%.

**Solution 7:**

- (c) Chlorine.

The compound formed which is strongly acidic in water, is HCl.



- (d) A dilute aqueous solution of hydrochloric acid gets gradually concentrated on distillation, till the concentration of the acid reaches 22.2% HCl by weight which boils at  $110^\circ\text{C}$ . When this concentration is reached, no further increase in concentration of the acid becomes possible by boiling. This is because vapours evolved before  $110^\circ\text{C}$  are vapours of water but at temperature above  $110^\circ\text{C}$  vapours consist mostly of molecules of HCl.

**Question 8:**

Hydrochloric acid contains (i) hydrogen (ii) chlorine. Prove it. Write equations for the reactions.

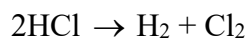
**Solution 8:**

We can prove that hydrochloric acid contains both hydrogen and chlorine by the following experiment.

Take a voltameter used for electrolysis of water, fitted with platinum cathode and graphite anode. Into the voltameter pour 4 molar HCl and pass direct current.

It is seen that a colourless gas is evolved at cathode and a greenish gas is evolved at anode. When a burning splinter is brought near a colourless gas, it bursts into flame thereby proving that it is hydrogen gas.

When moist starch iodide paper is held in the greenish yellow gas, it turns blue black, thereby proving that the gas is chlorine.



This experiment proves that hydrochloric acid contains both hydrogen and chlorine.

### Question 9:

Name:

- (a) black metallic oxide which reacts with hydrochloric acid to give a coloured solution.
- (b) two colourless gases, which when mixed produce a white solid.
- (c) two gases which chemically combine to form a liquid
- (d) a chloride which is soluble in excess of ammonium hydroxide
- (e) the chemical in which gold can be dissolved
- (f) the experiment which demonstrates that hydrogen chloride is soluble in water.
- (g) the gas produced when chlorine water is exposed to sunlight

### Solution 9:

- (a) Manganese dioxide
- (b) Hydrogen chloride and ammonia
- (c) Hydrogen and oxygen
- (d) AgCl (Silver chloride)
- (e) Aqua regia
- (f) Fountain experiment
- (g) Hydrogen chloride gas

### Question 10:

Give reasons for the following:

- (a) An aqueous solution of chlorine is acidic in nature

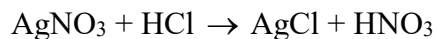
Hint :  $\text{Cl}_2 + \text{H}_2\text{O} \rightarrow \text{HCl} + \text{HClO}$

- (b) silver nitrate solution can be used to distinguish HCl from  $\text{HNO}_3$

### Solution 10:

- (a) An aqueous solution of chlorine is acidic as it dissolves in water to form hydrochloric and hypochlorous acids.

- (b) Silver nitrate reacts with hydrochloric acid to form thick curdy white ppt. of silver chloride whereas silver nitrate does not react with nitric acid.



(White ppt.)

**Question 11:**

Solution A reacts with an acid B (which gives greenish yellow gas on reacting with oxidizing agents like  $\text{Pb}_3\text{O}_4$  to give white precipitate C insoluble in nitric acid but soluble in ammonium hydroxide. Name A, B and C.

**Solution 11:**

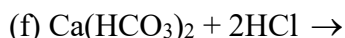
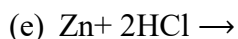
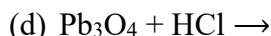
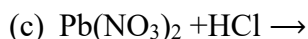
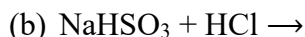
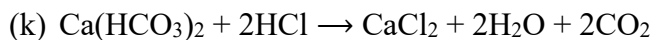
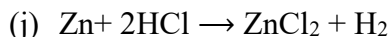
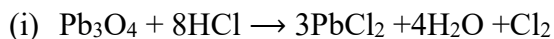
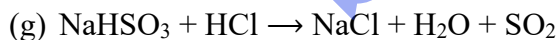
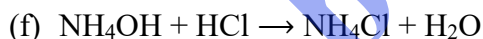
A is Silver nitrate

B is Hydrochloric acid

C is Silver chloride

**Question 12:**

Complete the following reactions and balance them.

**Solution 12:**

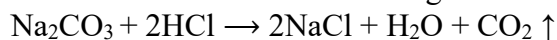
**Question 13:**

How will the action of dilute hydrochloric acid enable you to distinguish between the following:

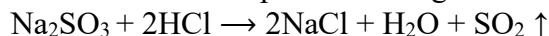
- (a) Sodium carbonate and sodium sulphite
- (b) sodium thiosulphate and sodium sulphite

**Solution 13:**

- a. Sodium carbonate on treating with dil. HCl results in the formation of sodium chloride with the liberation of carbon dioxide gas.



Sodium sulphite on treating with dil. HCl results in the formation of sodium chloride with the liberation of sulphur dioxide gas.



- b. Sodium thiosulphate reacts with dil. HCl to produce sulphur dioxide gas and precipitates yellow sulphur.



Sulphur is not precipitated when sulphites are treated with dil. HCl.

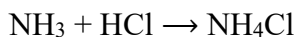
**Question 14:**

Give three distinct test [apart from using an indicator] you would carry out with solution of HCl to illustrate the typical properties of an acid.

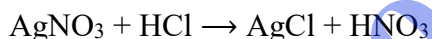
**Solution 14:**

Three tests are:

HCl gas gives thick white fumes of ammonium chloride when glass rod dipped in ammonia solution is held near the vapours of the acid.



With silver nitrate HCl gives white precipitate of silver chloride. The precipitate is insoluble in nitric acid but soluble in ammonium hydroxide.



A greenish yellow gas is liberated when concentrated hydrochloric acid is heated with oxidizing agent like manganese dioxide.

**Question 15:**

$\text{MnO}_2$ ,  $\text{PbO}_2$  and red lead react with conc. HCl acid liberates  $\text{Cl}_2$ .

What is the common property being shown by these metal oxides?

**Solution 15:**

$\text{MnO}_2$ ,  $\text{PbO}_2$  and red lead react with conc. HCl acid to liberate  $\text{Cl}_2$ . This shows that hydrochloric acid is oxidized to chlorine by oxidizing agents.





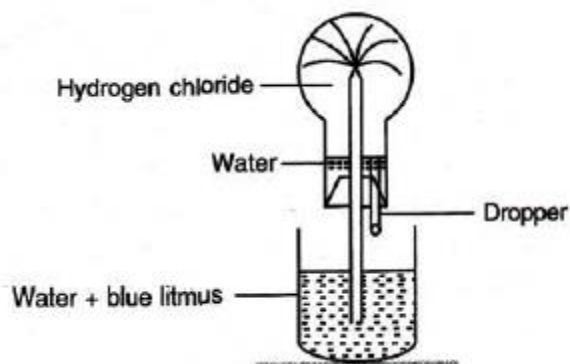


**Question (2005) 22:**

(a) Write balanced equations.

- (i) copper oxide and dilute hydrochloric acid.
- (ii) Manganese (iv) oxide and concentrated hydrochloric acid.

(b) (i) Name the experiment illustrated below.

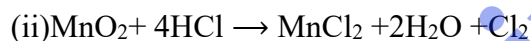
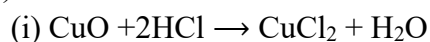


(ii) Which property of hydrogen chloride is demonstrated by this experiment?

(iii) State the colour of the water that has entered the round bottomed flask.

**Solution (2005) 22:**

(a)



(b)

(i) The experiment is called Fountain Experiment.

(ii) This experiment shows that hydrogen chloride is highly soluble in water.

(iii) Red

**Question (2007) 23:**

Write balanced equations for the reaction of dilute hydrochloric acid with each of the following:

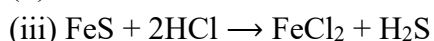
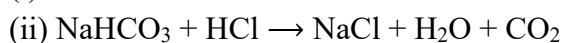
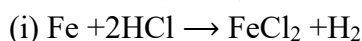
(a) iron

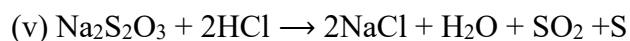
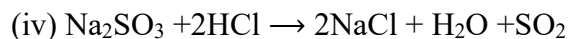
(b) sodium hydrogencarbonate,

(c) iron (II) sulphide

(d) sodium sulphite

(e) sodium thiosulphate solution

**Solution (2007) 23:**

**Question (2008) 24:**

What property of hydrogen chloride is demonstrated when it is collected by downward delivery (upward displacement)?

**Solution (2008) 24:**

When hydrogen chloride is collected by downward delivery or upward displacement, it shows that it is heavier than air.

**Question (2008) 25:**

Why is hydrogen chloride not collected over water?

**Solution (2008) 25:**

Hydrogen chloride is not collected over water as it is soluble in water.

**Question (2008) 26:**

Write the equations for the following reactions:

(a) Dilute hydrochloric acid and sodium thiosulphate,

(b) Dilute hydrochloric acid and lead nitrate solution.

**Solution (2008) 26:**