

## Atomic Structure

### Points to Remember:

- Anything which occupies space and has a mass is called matter.
- Smallest particle of matter is called the atom.
- In 1808 John Dalton, an English Scientist, described the atom as the smallest particle exhibiting all the properties of a particular element. According to him, atom could not be divided further into smaller particles.
- Later discoveries led to modifications of Dalton's Theory and now is known that an atom consists of three-fundamental particles; electrons, protons, neutrons.
- An atom is electrically neutral i.e. The number of protons in an atom equal to the number of electrons in it.
- The distribution of electrons in the various orbits of an element is called the electronic configuration of that element. The maximum number of electrons in a shell is given by the formula  $2n^2$ . Where n is the serial number of the shell. But the outer most shell cannot have more than 8 electrons.
- The noble gases do not react chemically because they have a stable configuration of 8 electrons in their outermost orbits.
- Elements are chemically reactive because they can gain or loose electrons to attain stable electronic configuration.
- The number of electrons donated or accepted by an atom during chemical combination is called its valency.
- The charged atoms formed due to the transfer of electrons are called ions. The positively charges ion is called a cation negatively charges ion is called an anion.
- The atoms of an element having the same atomic number but different mass number are called the isotopes of element.
- The number of protons in the nucleus of an atom is called the atomic number, which is equal to the number of electrons in a neutral atom.
- The- sum of the number of protons and the number of neutrons in the nucleus of an atom is called its mass number.

### Exercise

#### 1. Fill in the blanks.

- (a) Dalton **said** that atoms could not be divided
- (b) An ion which has a positive charge is called a **cation**.
- (c) The outermost shell of an atom is known as **valence shell**.
- (d) The **nucleus** of an atom is very hard and dense.
- (e) Neutrons are **neutral** particles having mass equal to that of protons.
- (f) Isotopes are the atoms of **an** element having the same atomic number but a different mass number.

**2. Write 'true' or 'false' for the following statements:**

(a) An atom on the whole has a positive charge.

**false**

(b) The maximum number of electrons in the first shell can be 8.

**false**

(c) The central part of the atom is called nucleus.

**True.**

**3. Give the following a suitable word/phrase.**

(a) The sub-atomic particle with negative charge and negligible mass.

(b) Protons and neutrons present in the nucleus.

(c) The electrons present in the outermost shell.

(d) Arrangement of electrons in the shells of an atom.

(e) The number of protons present in the nucleus of an atom.

(f) The sum of the number of protons and neutrons of an atom.

(g) Atoms of same element with same atomic number but a different mass number.

(h) The smallest unit of an element which takes part in a chemical reaction.

**Answer:**

(a) Neutron

(b) Mass number

(c) Valency

(d) Orbits or Valence shells

(e) Atomic number

(f) Mass number

(g) Isotopes

(h) Atom

**4. Multiple Choice Questions**

**(a) The outermost shell of an atom is known as**

1. valency

2. **valence electrons**

3. nucleus

4. valence shell

**(b) The number of valence electrons present in magnesium is**

1. **two**

2. three
3. four
4. five

(c) The sub atomic particle with negative charge is

1. proton
2. neutron
3. **electron**
4. nucleon

(d) If the atomic number of an atom is 17 and mass number is 35 then number of neutron will be

1. 35
2. 17
3. **18**
4. 52

(e) The number of electrons in an atom is equal to number of

1. **protons in a neutral atom**
2. neutrons in a neutral atom
3. nucleons in a neutral atom
4. none of the above

(f) The sum of number of protons and number of neutrons present in the nucleus of an atom is called its

1. **mass number**
2. atomic number
3. number of electrons
4. all of the above

**Question 5.**

Name three fundamental particles of the atom. Give the symbol with charge, on each particle.

**Answer:**

The fundamental particles of the atom are: electrons, protons and neutrons.

Particle	Symbol	Charge
electron	$e^-$	-1 or $1.602 \times 10^{-19}$ C. Where -1 represent its one unit negative electrical charge
proton	$p^+$	+ 1 or $1.602 \times 10^{-19}$ C. Where +1 represents one unit +ve electrical charge.
neutron	$n^0$	0

**Question 6.**

Define the following terms:

- (a) Atomic number
- (b) Mass number
- (c) Nucleons
- (d) Valence shell

**Answer:**

**(a) Atomic number:** Atomic number refers to the number of protons present in an atom. It is denoted by Z. Example: An atom of oxygen contains 8 proton Therefore its atomic number is 8.

**(b) Mass number:** Mass number refers to the sum of the number of protons and neutrons present in the nucleus of an atom and denoted by A Mass number = Number of protons + Number of neutrons.

**(c) Nucleons:** The protons and neutrons collectively are known as nucleons.

**(d) Valence Shell:** The outermost shell of an atom is known as its valence shell.

### Question 7.

Mention briefly the salient features of Dalton's atomic theory (five points).

**Answer:**

**Salient features of Dalton's atomic theory:**

1. Matter consists of very small and indivisible particles called atoms, which can neither be created nor can be destroyed.
2. The atoms of an element are alike in all respects i.e. size, mass, density, chemical properties but they differ from the atoms of other elements.
3. Atoms of an element combine in small numbers to form molecules of the element.
4. Atoms of one element combine with atoms of another element in simple whole number ratio to form molecules of compounds.
5. Atoms are the smallest units of matter that can take part in a chemical reaction during which only rearrangement of atoms takes place.

### Question 8.

(a) What are the two main features of Rutherford's atomic model?

(b) State its one drawback.

**Answer:**

(a) According to Rutherford's model an atom consists of:

1. **The centrally located nucleus:** The nucleus is a centrally located positively charged mass. The entire mass of the atom is concentrated in it. It is the densest part of the atom. Its size is very small as compare to the atom as a whole.
2. **The outer circular orbits:** Electrons revolve in circular orbits (shell) in the space available around the nucleus. An atom is electrically neutral i.e., number of protons and electrons present in an atom are equal.

(b) Rutherford's atomic model could not explain the stability of the atom as it is like a solar system, the sun is at the centre and the planets revolve around it, in an atom the electrons revolve around the centrally located nucleus containing protons.

### Question 9.

What are the observations of the experiment done by Rutherford in order to determine the structure of an atom?

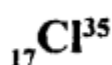
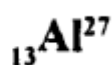
**Answer:**

Following were the observations made by Rutherford:

1. Most of the alpha particle passed straight through the foil without any deflection from their path.
2. A small fraction of them were deflected from their original path by small angles.
3. Only a few particles bounced back.

### Question 10.

State the mass number, the atomic number, number of neutrons and electronic configuration of the following atoms.

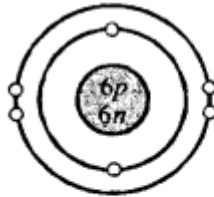


Also, draw atomic diagrams for them.

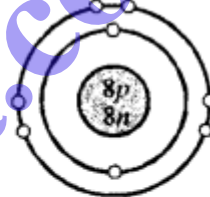
Answer:

Name of elements	Atomic number	Atomic Mass	No. of proton	No. of electrons	No. of neutrons	Electronic configuration
${}^1_6\text{C}$	6	12	6	6	6	2, 4
${}^{16}_8\text{O}$	8	16	8	8	8	2, 6
${}^{19}_9\text{F}$	9	19	9	9	10	2, 7
${}^{20}_{10}\text{Ne}$	10	20	10	10	10	2, 8
${}^{27}_{13}\text{Al}$	13	27	13	13	14	2, 8, 3
${}^{35}_{17}\text{Cl}$	17	35	17	17	18	2, 8, 7

Carbon  ${}^{12}_6\text{C}$



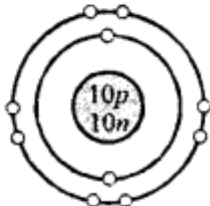
Oxygen  ${}^{16}_8\text{O}$



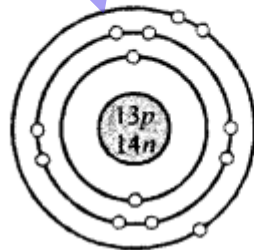
Fluorine  ${}^{19}_9\text{F}$



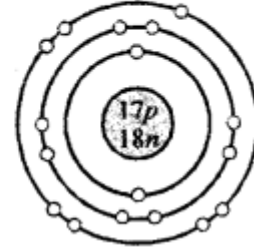
Neon  ${}^{20}_{10}\text{Ne}$



Aluminium  ${}^{27}_{13}\text{Al}$



Aluminium  ${}^{35}_{17}\text{Cl}$



**Question 11.**

What is variable valency? Name two elements having variable valency and state their valencies.

**Answer:**

Variable valency: Some elements exhibit more than one valency. They are said to have variable valency, e.g. Iron, copper, tin, lead.

Iron Fe Fe<sup>2+</sup> or Fe<sup>3+</sup>

Copper cu cu<sup>+</sup> or cu<sup>2+</sup>

**Question 12.**

The atomic number and the mass number of sodium are 11 and 23 respectively. What information is conveyed by this statement.

**Answer:**

Atomic number = 11; No of protons = 11

Mass number = 23 = Number of protons + Number of neutrons.

No of neutrons = 23-11 = 12.

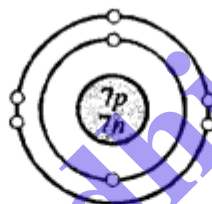
**Question 13.**

Draw the diagrams representing the atomic structures of the following:

(a) Nitrogen (b) Neon

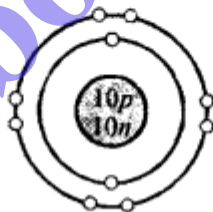
**Answer:**

(a) Nitrogen  ${}^{14}_7\text{N}$



Electronic configuration  
2, 5

(b) Neon  ${}^{20}_{10}\text{Ne}$



Electronic configuration  
2, 8

**Question 14.**

Explain the rule with example according to which electrons are filled in various energy levels,

**Answer:**

The maximum number of electrons that can be present in any shell or orbit of an atom is given by the formula  $2n^2$ , where  $n$  is the serial number of the shell.

Therefore:

K shell,  $n = 1$ , no. of electrons =  $2 \times 1^2 = 2$

L shell,  $n = 2$ , no. of electrons =  $2 \times 2^2 = 8$

M shell,  $n = 3$ , no. of electrons =  $2 \times 3^2 = 18$

N shell,  $n = 4$ , no. of electrons =  $2 \times 4^2 = 32$

Electrons are not accommodated in a given shell, unless the inner shells are filled.

That is, the shells are filled in a stepwise manner.

**Example:**

Elements	Atomic number	Electronic Configuration			
		K	L	M	N
Helium	2	2			
Boron	5	2	3		
Oxygen	8	2	6		
Neon	10	2	8		

**Question 15.**

The atom of an element is made up of 4 protons, 5 neutrons and 4 electrons. What is its atomic number and mass number?

**Answer:**

Protons = 4, neutrons = 5, electrons = 4

Atomic number = 4,

Mass number =  $4 + 5 = 9$

**Question 16.**

(a) What are the two main parts of which an atom is made of?

(b) Where is the nucleus of an atom situated ?

(c) What are orbits or shells of an atom ?

**Answer:**

(a)

1. The centrally located nucleus
2. The outer circular orbits.

(b) The nucleus is a centrally located positively charged mass.

(c) The circular orbits (shell present) in the space available around the nucleus on which electrons revolve are called orbits or shells of an atom.

**Question 17.**

What are isotopes? How does the existence of isotopes contradict Dalton's atomic theory?

**Answer:**

Atoms of an element must have the same atomic number, but their mass number can



be different due to the presence of different number of neutrons. These atoms of an element having different number of neutrons are called groups.

According to Dalton's theory, all atoms of an element are similar to all respects, for example, they have the same shape, size etc. and have similar physical and chemical properties like mass, density and reactivity. Whereas isotopes of an element have atoms that are similar as they have same number of protons and electrons but differ in the number of neutrons. So, the isotopes have atoms that are not similar in all aspects.

**Question 18.**

Complete the table below by identifying A, B, C, D, E and F.

**Answer:**

Element	Symbol	Number of protons	Number of neutrons	Number if electrons
Fluorine	${}^9\text{F}^{19}$	9	A	B
Aluminium	C	D	14	13
Potassium	${}_{19}\text{K}^{39}$	E	F	19

**Ans.**

Element	Symbol	Number of protons	Number of neutrons	Number if electrons
Fluorine	${}^9\text{F}^{19}$	9	10	9
Aluminium	${}_{13}\text{Al}^{27}$	13	14	13
Potassium	${}_{19}\text{K}^{39}$	19	20	19

**EXTRA PRACTICE QUESTIONS**

**FILL IN THE BLANKS**

1. Atomic theory of matter was given by **Democritus**.
2. Gases become good conductors of electricity under **low** (low/high) pressure and **high** (low/high) voltage electrical discharge.
3. Rutherford won the Nobel Prize in 1908 for his investigations in the field of **atomic theory**.
4. The electrons revolving in shells close to the nucleus are called **bound** electrons.
5. In an atom, the last shell containing electrons is called its **valence** shell.
6. The protons and neutrons are held firmly in the nucleus by strong **nuclear forces**.
7. Amount of energy stored inside the **nucleus** is known as nuclear energy.
8. The  $\gamma$ -rays are **electromagnetic** radiations, like light.
9. Nuclear energy can be used to make **electricity**.

## MULTIPLE CHOICE QUESTIONS

Tick the most appropriate answer.

1. Which of these have similar chemical properties but different set of physical properties and mass number?

1. isobars
2. elements
3. **isotopes**
4. none of these

2. Which one of the following scientists put forward the theory regarding the extra-nuclear structure of the atom?

1. James Chadwick
2. John Dalton
3. Rutherford
4. **Niels Bohr**

3. Which of the following scientists observed that cathode rays consists of negatively charged particles?

1. John Dalton
2. **J J Thomson**
3. James Chadwick
4. Democritus

4. In 1908, R A Millikan found the charge of an electron to be

1.  **$1.6 \times 10^{-19} \text{ C}$**
2.  $1.6 \times 10^{-15} \text{ C}$
3.  $1.4 \times 10^{-19} \text{ C}$
4.  $1.6 \times 10^{-16} \text{ C}$ .

5. How many times is a proton heavier than an electron?

1. 1500
2. 1800
3. **1840**
4. none of these

6. Who discovered that radioactive rays are given out by uranium?

1. Henri Becquerel
2. **Marie Curie**
3. Pierre Curie
4. Schmidt

7.  $\beta$ -particles are actually a stream of

1. protons
2. **electrons**
3. neutrons
4. none of these

8. Which of the following is used as a coolant in nuclear reactor?

1. **liquid sodium**
2. liquid nitrogen
3. cadmium steel rods
4. plutonium

**MATCH THE COLUMNS**

- |                          |                                       |
|--------------------------|---------------------------------------|
| <b>1. John Dalton</b>    | <b>a. Neutrons</b>                    |
| <b>2. Atom</b>           | <b>b. Model of the atom</b>           |
| <b>3. Chadwick</b>       | <b>c. The Atomic Theory of Matter</b> |
| <b>4. Rutherford</b>     | <b>d. Positively charged</b>          |
| <b>5. Protons</b>        | <b>e. Electrically neutral</b>        |
| <b>6. Nuclear fusion</b> | <b>f. Generator of electricity</b>    |
| <b>7. Gamma-rays</b>     | <b>g. High penetrating power</b>      |
|                          | <b>h. Energy in sun</b>               |

**Ans.**

- |                   |                                |
|-------------------|--------------------------------|
| 1. John Dalton    | c. The Atomic Theory of Matter |
| 2. Atom           | e. Electrically neutral        |
| 3. Chadwick       | a. Neutrons                    |
| 4. Rutherford     | b. Model of the atom           |
| 5. Protons        | d. Positively charged          |
| 6. Nuclear fusion | h. Energy in sun               |
| 7. Gamma-rays     | g. High penetrating power      |

## TRUE \ FALSE

1. Elements combine in definite proportions to form compounds.  
**True.**
2. Elements having a tendency to accept electrons are called metals.  
**False.** Element having a tendency to release/donate electrons are called metals.
3. Electrons are present in the nucleus.  
**False.** Electrons are present in the shells.
4. Neils Bohr gave the theory to explain the arrangement of protons in an atom.  
**False.** Neils Bohr gave the theory to explain the arrangement of electrons in an atom.
5. Cathode rays are a stream of fast moving particles that travel in random directions.  
**False.** Cathode rays are a stream of fast moving particles that travel in straight line.
6. Uranium-235 and uranium-238 both undergo nuclear fission.  
**True.**
7. After the emission of a beta-particle, the atomic number of the atom increases by one.  
**True.**

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## WRITE SHORT ANSWERS

### Question 1.

What is an atom?

#### Answer:

An atom is defined as the smallest particle of a molecule which is further subdivided into electrons, protons and neutrons.

### Question 2.

Describe a discharge tube.

#### Answer:

A discharge tube is a cylindrical glass tube in which a gas can be filled and which is fitted with metal plates called electrodes at each end.

### Question 3.

What is the central part of an atom known as?

#### Answer:

Nucleus consisting of neutrons and protons.

### Question 4.

What do you mean by electron transition?

#### Answer:

The jumping of an electron from one energy level to another is called electronic transition.

### Question 5.

Define an ionic compound.

#### Answer:

An ionic compound is defined as the compound formed by the oppositely charged ions which attract each other to form a chemical bond e.g. Sodium chloride.

### Question 6.

Define the atomic mass number of an atom.

#### Answer:

The atomic mass number refers to total number of protons and neutrons in a nucleus.

### Question 7.

Name the isotopes of hydrogen.

#### Answer:

There are 3 (three) isotopes of hydrogen

1. Protium  ${}_1\text{H}^1$
2. Deuterium  ${}_1\text{H}^2$
3. Tritium  ${}_1\text{H}^3$

**Question 8.**

What is radioactivity?

**Answer:**

Radioactivity is the phenomena in which certain substances emit radiations spontaneously.

**Question 9.**

Name three inert gases.

**Answer:**

Inert gas: Argon, Helium, Neon.

**Question 10.**

What is the charge on a neutron ?

**Answer:**

No charge.

**Question 11.**

What is the number of neutrons present in a potassium atom ?

**Answer:**

20.

**ANSWER IN DETAIL****Question 1.**

State the postulates of the Dalton's atomic theory.

**Answer:****Postulates of Dalton's atomic theory:**

1. All matter is made up of minute particles called atoms.
2. Atom are indivisible. They can neither be created nor destroyed.
3. All atoms of a given element are identical in mass and size.
4. The atoms of one element are different from the atoms of other elements.
5. During a chemical reaction, the atoms of an element combine with atoms of another element in a fixed ratio.

**Question 2.**

Name the three subatomic particles. How are they different from each other ?

**Answer:**

The three subatomic particles are:

1. **Electron:** It is represented by  $e$  it has a unit negative charge and possesses mass equal to  $1/1837$  times that of a proton.
2. **Proton:** It is represented by  ${}_{+1}p^1$ . It has a unit positive charge. Its mass is equal to the mass of a hydrogen atom.
3. **Neutron:** It is denoted as  ${}_0n^1$ . It has no charge on it. Its mass is almost equal to the mass of hydrogen atom.

### Question 3.

Enlist the features of Bohr's model of an atom. How did he correct the Rutherford's model ?

**Answer:**

**Bohr's Postulates:**

1. The extra-nuclear region of an atom consists of electrons revolving in particular orbits called shells.
  2. Each shell has a fixed energy level.
  3. As long as an electron revolves in the same shell, the energy remains constant. During electronic transition energy is either released or absorbed.
- Bohr corrected Rutherford's model by suggesting the concepts of energy levels and electronic transition.

### Question 4.

What is the maximum number of electrons that L shell can accommodate ? Give reasons for your answer.

**Answer:**

The maximum possible number of electrons that a shell can accommodate can be given by the formula  $2n^2$ .

(Here n is the number of shell)

In case of L shell  $n = 2$

$$\therefore 2n^2 = 2(2)^2 = 8$$

Thus L shell can accommodate maximum number of 8 electrons.

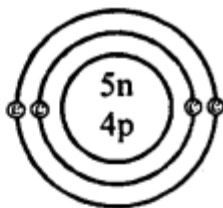
### Question 5.

Draw the electronic configuration of the following elements clearly stating the number of neutrons and protons in their nucleus.

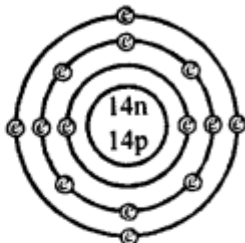
- a. Beryllium
- b. Silicon
- c. Chlorine
- d. Aluminium
- e. Nitrogen

**Answer:**

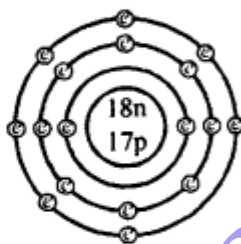
a. Beryllium



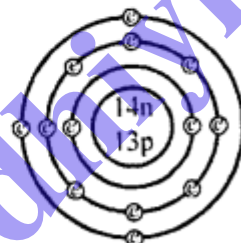
b. Silicon



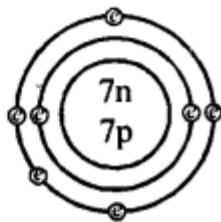
c. Chlorine



d. Aluminium



e. Nitrogen



### Question 6.

How are isobars different from isotopes ?

**Answer:**

**Isotopes** are defined as elements which have the same atomic number but different mass number e.g. hydrogen has isotopes namely.



1. Protium  ${}_1\text{H}^1$
2. Deuterium  ${}_1\text{H}^2$
3. Tritium  ${}_1\text{H}^3$

**Isobars** are the elements which have same mass number but different atomic number e.g. potassium and argon are isobars.

### Question 7.

Write a note on valence shell.

#### Answer:

The outermost shell of an atom is called its valence shell. The number of electrons that an atom needs to accept, donate or share so as to form a complete outer octet (eight electrons) is called the valency of that element. So, the valency of an element can be defined as its combining capacity with other elements. The valency of an element is always a whole number.

### Question 8.

What is nuclear fission ? How is it different from nuclear fusion ?

#### Answer:

**Nuclear fission:** The process in which a heavy nucleus, when bombarded with slow moving neutrons, splits into two or more lighter nuclei accompanied with a release of tremendous amount of energy.

**Nuclear fusion:** The process in which two light nuclei combine together at very high temperature, and form a stable and a heavier nucleus and is accompanied with a release of tremendous amount of energy.

### Question 9.

What are the various components of a nuclear power plant?

#### Answer:

Nuclear power plant has the enlisted components.

1. **Fuel:** Uranium – 235 or plutonium – 239 which undergo fission easily are used as fuels.
2. **Moderator:** Heavy water or rods of graphite are used as moderators. The controlling rods made of cadmium or boron steel are used to regulate the number of neutrons in the reactor.
3. **Coolant:** As a great amount of heat is released during this process, a coolant is also required. Liquid sodium and heavy water are generally used as coolants.

### Question 10.

How are radioactive radiations useful in medicines ?

#### Answer:

Controlled radiation is used during treatment of cancer, as they can kill cancer cells. Cobalt-60 is used for destroying malignant cells in patients suffering from cancer. Iodine-131 is used for studying disorders of thyroid gland. Sodium-24 is used for examining the circulation of blood.